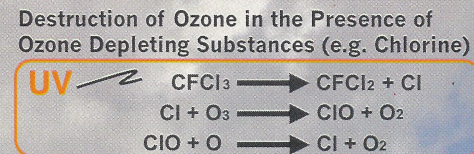
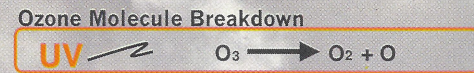
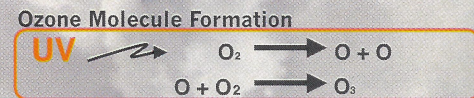
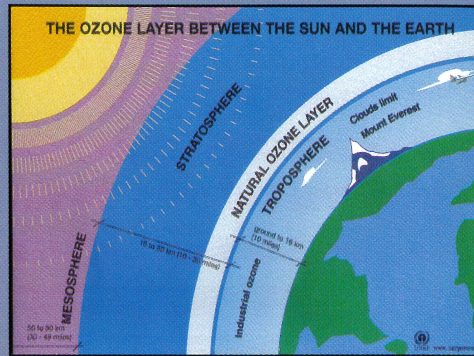


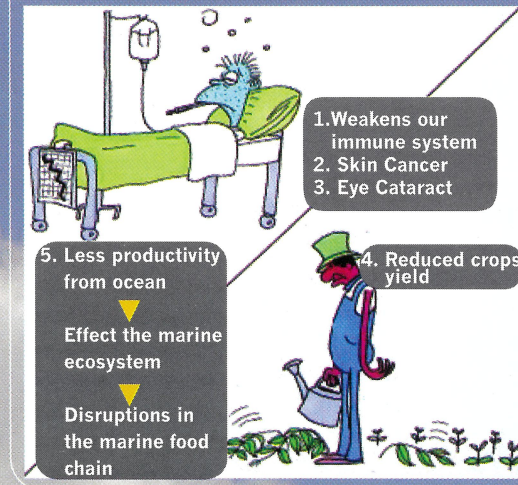
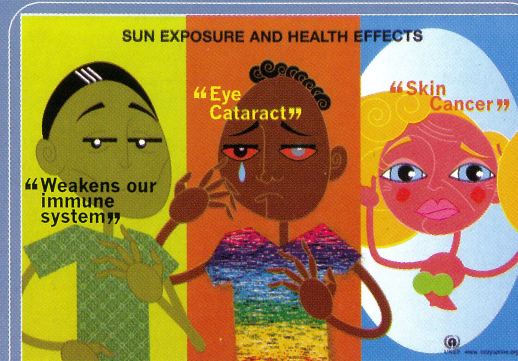


DO YOU KNOW...

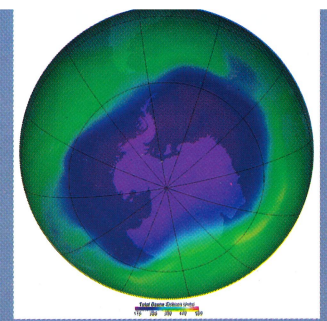
Our earth is protected by a thin ozone layer which filters out the harmful ultraviolet (UV) radiation. Ozone is a molecule that contains three atoms of oxygen and thus has the formula O₃. Ozone occurs naturally in the Earth's upper atmosphere – 10 to 20 miles above the Earth's surface where it forms a protective layer that shields us from the sun's harmful ultraviolet rays. This beneficial ozone is gradually being destroyed by manmade chemicals. For people, overexposure to UV rays can lead to skin cancer, cataracts and weakened immune systems. Increased UV can also lead to reduced crop yield and disruptions in the marine food chain.



EFFECTS ON OZONE LAYER DEPLETION



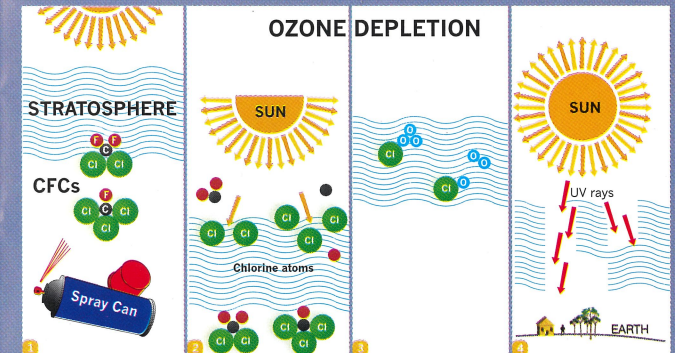
What's CFC?



CFC or Chlorofluorocarbon is a chemical compound that contains atoms of carbon, fluorine and chlorine. CFCs can deplete the ozone layer and reduce its ability to protect against the ultraviolet radiation consequently. The increase of ultraviolet radiation from the sun will render serious negative effects to human and ecosystem such as increase cases of skin cancer, cataract and reduced immunity. CFCs was the major cause of ozone depletion in the upper atmosphere.

Hydrochlorofluorocarbons (HCFCs) are similar to CFCs but less destructive to ozone. They are used as transition replacements for CFCs, but are to be phased out by the year 2030.

Hydrochlorofluorocarbons (HCFCs) do not have any potential for the destruction of ozone/ozone friendly.



CFCs are released in the air and travel up to the stratosphere.

CFCs are hit and broken by the sun's UV rays in the stratosphere. Chlorine atoms are released.

The chlorine atoms hit and break the ozone molecules that form our protective ozone layer. A chlorine atom can spend a hundred years breaking ozone molecules in the stratosphere.

While the ozone layer is depleted, more UV rays can go through and harm us.

